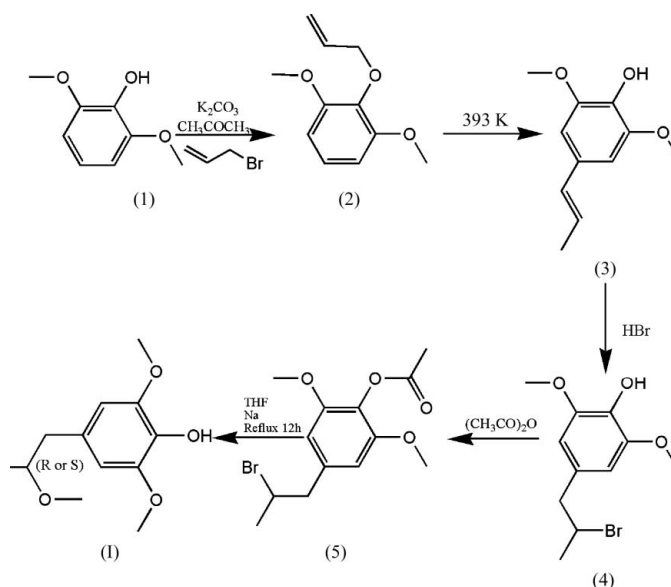


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Key indicators

Single-crystal X-ray study
 $T = 292\text{ K}$
Mean $\sigma(\text{C}-\text{C}) = 0.007\text{ \AA}$
 R factor = 0.057
 wR factor = 0.164
Data-to-parameter ratio = 8.3For details of how these key indicators were
automatically derived from the article, see
<http://journals.iucr.org/e>.**(*RS*)-2,6-Dimethoxy-4-(2-methoxypropyl)phenol**The title compound, $\text{C}_{12}\text{H}_{18}\text{O}_4$, crystallizes with two molecules
in the asymmetric unit. The structure features $\text{O}-\text{H}\cdots\text{O}$
hydrogen-bonded double chains.Received 15 August 2005
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Comment

The mechanism of phenol oxidation by tyrosinase has been
well studied (Krol & Bolton, 1997). However, little work has
been done to determine the influence of substituents on the
reaction. Recently, we have synthesized the title compound,
(I), and its crystal structure is reported here.

Compound (I) crystallizes in the non-centrosymmetric space group Pc , with two independent molecules in the asymmetric unit. In the molecular structure (Fig. 1 and Table 1), the two molecules have similar geometric parameters about the two chiral centres, C10 and C22; the bond lengths and angles are unremarkable. The crystal structure is stabilized by intermolecular $\text{O}-\text{H}\cdots\text{O}$ hydrogen bonds, with the phenol groups acting as donors and the ether O atoms as acceptors (Table 2). These hydrogen bonds link the molecules into double chains (Fig. 2).

Experimental

2,6-Dimethoxyphenol, (1), was synthesized according to the literature procedure of Wu *et al.* (1997). To a basic acetone solution of 2,6-dimethoxyphenol (0.35 g, 2 mmol) was added K_2CO_3 (0.14 g); a solution of 3-bromoprop-1-ene (15 ml) in acetone was then added dropwise. The mixture was stirred at 308 K for 5 h, filtered and

purified by column chromatography on silical gel, using petroleum-ethyl acetate (3:2 v/v), to give compound (2) (0.30 g, 1.5 mmol). Product (2) (0.19 g, 1 mmol) in tetrahydrofuran (THF, 20 ml) was refluxed for 1 h at 393 K and then cooled to room temperature. To the solution was added hydrobromic acid (5 ml, 6 M) and the mixture stirred for 1 h. Removal of the solvent under reduced pressure and subsequent purification by flash chromatography (using acetone as eluent) afforded compound (4) (0.15 g, 0.5 mmol). The hydroxyl group of compound (4) was protected with acetic anhydride, yielding compound (5). To a solution of (5) in THF (20 ml) was added sodium (0.07 g) and the mixture was refluxed for 12 h at 323 K to afford the title compound, (I). Crystals suitable for X-ray analysis were obtained by recrystallization from ethanol.

Crystal data

$C_{12}H_{18}O_4$	$D_x = 1.190 \text{ Mg m}^{-3}$
$M_r = 226.26$	Mo $K\alpha$ radiation
Monoclinic, Pc	Cell parameters from 2536 reflections
$a = 13.740 (3) \text{ \AA}$	$\theta = 2.6\text{--}25.4^\circ$
$b = 9.405 (2) \text{ \AA}$	$\mu = 0.09 \text{ mm}^{-1}$
$c = 9.997 (2) \text{ \AA}$	$T = 292 (2) \text{ K}$
$\beta = 102.190 (4)^\circ$	Block, colorless
$V = 1262.8 (5) \text{ \AA}^3$	$0.30 \times 0.20 \times 0.10 \text{ mm}$
$Z = 4$	

Data collection

Bruker SMART APEX CCD area-detector diffractometer	1809 reflections with $I > 2\sigma(I)$
ω scans	$R_{\text{int}} = 0.081$
Absorption correction: none	$\theta_{\text{max}} = 26.0^\circ$
6654 measured reflections	$h = -16 \rightarrow 16$
2481 independent reflections	$k = -11 \rightarrow 6$
	$l = -12 \rightarrow 11$

Refinement

Refinement on F^2	H-atom parameters constrained
$R[F^2 > 2\sigma(F^2)] = 0.057$	$w = 1/[\sigma^2(F_o^2) + (0.1154P)^2]$
$wR(F^2) = 0.164$	where $P = (F_o^2 + 2F_c^2)/3$
$S = 0.97$	$(\Delta/\sigma)_{\text{max}} < 0.001$
2481 reflections	$\Delta\rho_{\text{max}} = 0.37 \text{ e \AA}^{-3}$
299 parameters	$\Delta\rho_{\text{min}} = -0.19 \text{ e \AA}^{-3}$

Table 1

Selected geometric parameters (\AA , $^\circ$).

C9–C10	1.502 (6)	C21–C22	1.501 (8)
C10–O4	1.437 (5)	C22–O8	1.431 (6)
C10–C11	1.491 (7)	C22–C23	1.555 (10)
O4–C10–C11	111.5 (4)	O8–C22–C21	105.3 (4)
O4–C10–C9	105.4 (3)	O8–C22–C23	112.4 (5)
C11–C10–C9	113.1 (4)	C21–C22–C23	111.8 (5)

Table 2

Hydrogen-bond geometry (\AA , $^\circ$).

$D\text{--}H\cdots A$	$D\text{--}H$	$H\cdots A$	$D\cdots A$	$D\text{--}H\cdots A$
O2–H2 \cdots O1	0.82	2.26	2.676 (4)	112
O2–H2 \cdots O8 ⁱ	0.82	2.05	2.774 (5)	148
O6–H6A \cdots O5	0.82	2.25	2.661 (4)	111
O6–H6A \cdots O4 ⁱⁱ	0.82	2.05	2.823 (4)	156

Symmetry codes: (i) $x + 1, y, z$; (ii) $x, y - 1, z + 1$.

H atoms were positioned geometrically (methyl C–H = 0.96 \AA , methylene C–H = 0.97 \AA , methine C–H = 0.98 \AA , aromatic C–H = 0.93 \AA and hydroxy O–H = 0.82 \AA) and included in the refinement in the riding-model approximation, with $U_{\text{iso}}(\text{H}) = xU_{\text{eq}}(\text{carrier})$

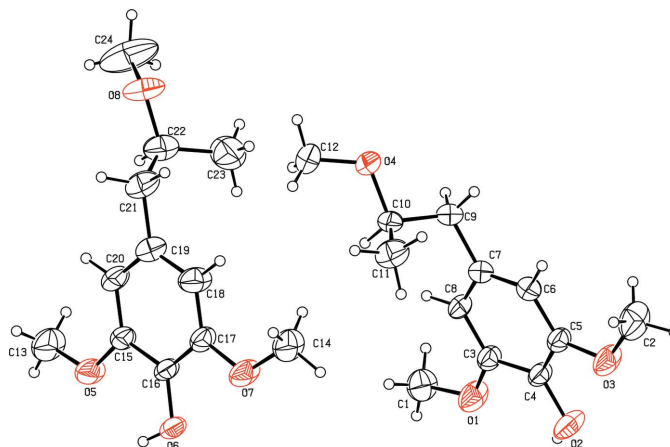


Figure 1

The molecular structure of the asymmetric unit of the title compound, showing 30% probability displacement ellipsoids.

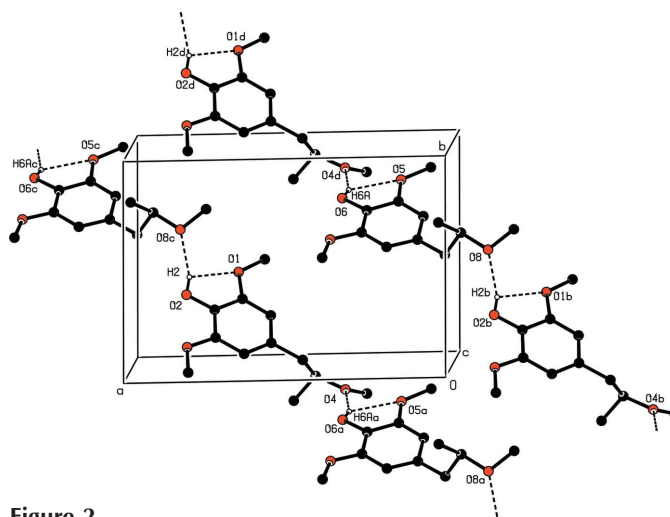


Figure 2

Plot of the crystal packing, showing the linkage of molecules by O–H \cdots O hydrogen bonding (dashed lines). Labels a, b, c, d correspond to the symmetry positions $(x, -1 + y, z)$, $(-1 + x, y, -1 + z)$, $(x + 1, y, z + 1)$ and $(x, y + 1, z)$, respectively.

atom), where $x = 1.5$ for methyl and hydroxy, and $x = 1.2$ for the others. In the absence of significant anomalous dispersion effects, Friedel pairs were averaged.

Data collection: *SMART* (Bruker, 2001); cell refinement: *SAINTE-Plus* (Bruker, 2001); data reduction: *SAINTE-Plus*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 1997); program(s) used to refine structure: *SHELXL97* (Sheldrick, 1997); molecular graphics: *PLATON* (Spek, 2003); software used to prepare material for publication: *PLATON*.

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